

June 2003

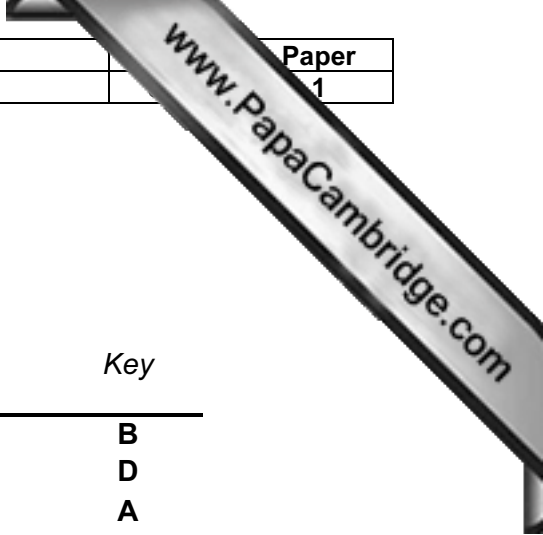
INTERNATIONAL GCSE

MARKING SCHEME

MAXIMUM MARK: 40

SYLLABUS/COMPONENT: 0620/01

CHEMISTRY
(Multiple Choice)



<i>Question Number</i>	<i>Key</i>	<i>Question Number</i>	<i>Key</i>
1	C	21	B
2	B	22	D
3	A	23	A
4	D	24	B
5	A	25	D
6	C	26	B
7	A	27	D
8	A	28	D
9	B	29	D
10	C	30	B
11	B	31	D
12	D	32	D
13	C	33	A
14	D	34	A
15	B	35	B
16	C	36	A
17	A	37	A
18	C	38	B
19	A	39	C
20	C	40	C

TOTAL 40

June 2003

INTERNATIONAL GCSE

MARKING SCHEME

MAXIMUM MARK: 80

SYLLABUS/COMPONENT: 0620/02

CHEMISTRY
(Core Paper 2)

- 1 (a) (i) Fe/Cu ALLOW Zn
(ii) C/N/S/F/C/Br
(iii) O/S
(iv) C
(v) Li/Na/K ALLOW F
(vi) CU/Zn/Br/Kr
- (b) argon - light bulbs;
chlorine - kills bacteria;
carbon - as lubricant;
helium - in balloons
- (c) (i) covalent
(ii) BrF₅ ALLOW F₅Br
(iii) ions/charged particles;
NOT: particles
not free to move in solid/free to move in molten/liquid state
- 2 (a) drop small tube in acid/loosen string/idea of mixing zinc and acid/let go of cotton
ALLOW: cut the string
NOT: heat (the acid)
NOT: pull the string
- (b) (i) correct plotting including 0-0 point (- 1 per omission or error)
(ii) best curve drawn and to go through origin
(iii) no more gas produced/reaction finished;
all zinc reacted/used up
- (c) graph drawn with faster initial rate and starting at 0-0;
ALLOW: straight line as initial rate
ends up at 55 cm³
- (d) (i) 2 (HCl)
(ii) zinc chloride
(iii) 136
IGNORE units
- (e) substance containing only one type of atom/substance which cannot be broken down to any other substance by chemical means
NOT 'can't be split' alone
NOT is a pure substance
- 3 (a) (i) evaporation/vaporisation/boiling
(ii) freezing/solidification
NOT: fusion
(iii) condensing/condensation/liquefaction
- (b) 2nd box ticked
- (c) A;
energy needed to overcome forces between molecules/idea of energy input/
taking in heat
- (d) (i) chlorine
(ii) bromine
(iii) sodium chloride

- (e) (i) diffusion
NOT: Brownian motion
- (ii) ammonium chloride
NOT: ammonia chloride
- (iii) ammonia diffuses or moves faster/HCl diffuses or moves slower/ammonia has lower mass/HCl higher mass/molecules of HCl and ammonia move at different speeds [1]
NOT: ammonia evaporates faster/HCl evaporates more slowly
- (f) neutralisation/acid base [1]
NOT: exothermic
NOT: addition
- (g) (i) thermometer [1]
(ii) reference to the solid or melting point of the solid is needed for the mark. boiling point of water too low to get solid to melt/boiling water cannot get to 155°C [1]
NOT: boiling point of water is only 100°C/boiling point of water too low.
NOT: water boils off first
- (iii) so that the liquid is the same temperature throughout/no hot or cold spots/so the tube is the same temperature as the thermometer/so heat can circulate in all places [1]
ALLOW: so that temperature of liquid is balanced
NOT: to keep temperature constant
- 4 (a) (i) breaking down of molecules substances using heat [1]
(ii) substance which speeds up a reaction [1]
NOT: alters/changes rate of reaction
NOT: speeds up and slows down rate
- (b) ethene/ethylene [1]
NOT: formula
- (c) (i) paraffin [1]
(ii) 4000g/4kg [1]
(correct unit needed)
(iii) C₂H₄; H₂ [2]
- (d) (i) two units polymerised with continuation bonds at either end and hydrogen atoms drawn [1]
ALLOW: -CH₂CH₂CH₂CH₂-
ALLOW: -[-CH₂CH₂-]_n
ALLOW: -[-CH₂-]_n
- (ii) addition (polymerisation) [1]
- 5 (a) (sodium) hydroxide/ammonia; → green/grey green; [2]
silver nitrate; → yellow; [2]
ALLOW: lead nitrate NOT: cream
ALLOW: bubble chlorine → grey/black (precipitate)
silver nitrate; → white; [2]
barium chloride/nitrate; → white; [2]
ALLOW: lead acetate

- (b) filtration/filtering or diagram of correct apparatus for filtration (filter paper be present on diagram)
 NOT: decanting
 sodium chloride through filter paper/shown on diagram;
 NOT: filtrate through filter paper
 evaporate off water from sodium chloride/suitable diagram [3]
 ALLOW: distilling off water
- (c) different atoms/elements
 (chemically) joined/bonded/combined (both points needed)
 (reference to mixtures = 0 unless qualified enough in time frame e.g. a mixture of elements which are then chemically combined) [1]
- (d) (i) chlorine/ Cl_2 [1]
 (ii) sodium/Na [1]
- 6 (a) potassium/magnesium/aluminium [1]
- (b) they did not have electricity/did not know about electrolysis/did not know the metal existed [1]
 NOT: did not have the right technology
- (c) (i) indication that bubbles produced rapidly or quickly/slower than magnesium but faster than zinc [1]
 OR number of bubbles produced intermediate between magnesium and zinc; uranium dissolved slower than magnesium but faster than zinc/dissolves at medium rate etc. [1]
 (ii) atoms of same element with different mass number/different number of neutrons/different nucleon number [1]
 NOT: compounds/molecules with different mass number
 (iii) indication of use for energy – nuclear power stations/nuclear energy [1]
 ALLOW: atomic/nuclear bombs
 NOT: curing cancer/medical uses
 NOT: 'for fuel'
- (d) magnesium oxide [1]
 ALLOW: MgO
- (e) (i) idea of mixture of (different) metals [1]
 (ii) alloys harder/stronger/decreased malleability/increased toughness/increased corrosion resistance/heat or electrical resistance increased [1]
 NOT: increase in melting point
 NOT: cheaper
 NOT: improving properties
- (f) removes oxygen from zinc oxide [1]
 ALLOW: definition of reduction involving oxidation numbers/electron transfer
- (g) (i) reversible reaction [1]
 ALLOW: equilibrium
 (ii) 76-80% [1]
- (h) (i) correct electronic structure of Mg (2.8.2) on diagram [1]
 (ii) loses two electrons/loses its valence electrons = 2
 forms Mg^{2+} ion = 1
 loses electron(s) = 1 [2]
 forms Mg^{2+} ion by losing electrons = 2

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INTERNATIONAL GCSE

MARKING SCHEME

MAXIMUM MARK: 80

SYLLABUS/COMPONENT: 0620/03

CHEMISTRY

(Extended Paper 3)

In the mark scheme if a word or phrase is underlined it (or an equivalent) is required for full marks of the mark.

(.....) is used to denote material that is not specifically required.

OR designates alternative and independent ways of gaining the marks for the question.

or indicates different ways of gaining the same mark.

COND indicates that the award of this mark is conditional upon a previous mark being gained.

- Unusual responses which include correct Chemistry that answers the question should always be rewarded-even if they are not mentioned in the marking scheme.
- All the candidate's work must show evidence of being marked by the examiner.

- 1 (a) A correct equation either CO or CO₂ as product
If not balanced but otherwise correct [1] ONLY [2]
- (b) (i) C + O₂ → CO₂ NOT word equation [1]
(ii) (higher in furnace) no oxygen left [1]
carbon dioxide reacts with carbon (to give carbon monoxide) [1]
- OR** incomplete combustion of carbon [2]
- OR** either equation gains both marks
CO₂ + C = 2CO or 2C + O₂ = 2CO
- OR** carbon dioxide reacts [1]
with carbon [1]
- (c) limestone + sand → slag [2]
OR calcium carbonate + silicon (IV) oxide → calcium silicate (+ carbon dioxide)
- For knowing that impurity is sand [1] ONLY
- Accept calcium oxide and silicon oxide
Accept lime
- (d) (i) Cutlery **or** chemical plant **or** watches **or** utensils **or** surgical instruments **or**
cars **or** sinks **or** aircraft **or** garden tools [1]
(ii) nickel **or** chromium **or** molybdenum **or** niobium **or** titanium [1]
(iii) blow air/oxygen through
carbon becomes carbon dioxide
carbon dioxide escapes as gas
silicon and phosphorus become oxides
calcium oxide or calcium carbonate
forms slag
Any FOUR NOT blast furnace [4]
- (e) anode tin NOT impure time [1]
cathode iron or steel [1]
tin salt **or** tin ions as electrolyte [1]
NOT oxide or hydroxide or carbonate

TOTAL = 16

- 2 (a) (i) 3 ignore any charges
(ii) high melting **or** boiling point
hard
poor conductor of electricity **or** heat
brittle
Any TWO [2]
NOT insoluble, dull, or malleable
- (iii) carbon, graphite diamond silicon, germanium [1]
silicon (IV) oxide **or** silica **or** silicon dioxide **or** silicon oxide
or sand **or** silicon carbide **or** named polymer [1]
- (iv) four around one [1]
cond looks tetrahedral **or** shows continuation [1]
For graphite layers [1] weak bonds between layers [1]
Accept any macromolecule, no link with (iii)
For polymer repeat unit [1] continuation [1]
- (b) (i) white precipitate [1]
COND upon a precipitate
dissolves in excess or forms solution [1]
- (ii) blue precipitate [1]
COND upon a precipitate
does not dissolve in excess [1]
- (c) (i) number of moles $\text{CO}_2 = 0.24/24 = 0.01$
conseq number of moles of CaCO_3 and $\text{MgCO}_3 = 0.01$
conseq number of moles of $\text{CaCO}_3 = 0.005$ [3]
- (ii) Calculate the volume of hydrochloric acid, 1.0 mole/dm^3 , needed to react with one tablet.
number of moles of CaCO_3 and MgCO_3 in one tablet = 0.01
Expect same as answer to (c)(i). NO marks to be awarded. Just mark sequentially to this response
conseq number of moles of HCl needed to react with one tablet = 0.02 [1]
- conseq** volume of hydrochloric acid, 1.0 mole/dm^3 , needed to react with one tablet = 0.02 dm^3 or 20 cm^3 [1]
- TOTAL = 16**
- 3 (a) (i) Correct equation [2]
For giving correct formula of alkane and alkene [1] only
Accept alkene and hydrogen
- (ii) chlorine [1]
COND light **or** 200°C **or** heat **or** lead tetraethyl
or high temperature MAX 1000°C [1]
ignore comment 'catalyst'
- (b) (i) same molecular formula [1]
different structures **or** structural formulae [1]
- (ii) but-2-ene or cyclobutane [1]
corresponding structural formula [1]
NOT 2-butene
- (c) butanol ignore numbers [1]
butane ignore numbers [1]
dibromobutane ignore numbers [1]

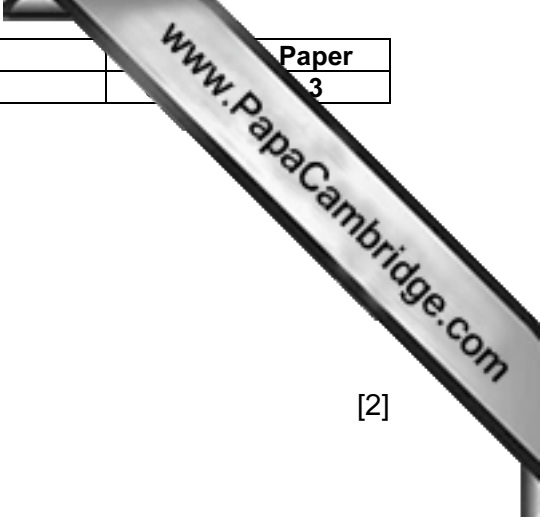
- (d) (i) propene
 $\text{CH}_3\text{—CH}=\text{CH}_2$
- (ii) Correct structure of repeat unit [1]
 ignore point of attachment of ester group
 COND upon repeat unit
 shows continuation [1]
 If chain through ester group [0] out of [2]
- (iii) do not decay or non-biodegradable
 shortage of sites or amount of waste per year
 visual pollution
 forms methane
 Any TWO [2]
- (iv) form poisonous **or** toxic gases **or** named gas CO, HCl HCN [1]
 NOT carbon dioxide, harmful, sulphur dioxide

TOTAL = 18

- 4 (a) (i) Correct equation [2]
 not balanced [1] ONLY
 $2\text{Pb}(\text{NO}_3)_2 = 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$
 $\text{Pb}(\text{NO}_3)_2 = \text{PO} + 2\text{NO}_2 + \frac{1}{2}\text{O}_2$
- (ii) potassium nitrate → potassium nitrite + oxygen [1]
- (b) (i) close **or** tightly packed [1]
 ordered **or** lattice [1]
 vibrational [1]
 NOT forces
- (ii) melting **or** freezing **or** fusion **or** solidification [1]
- (c) (i) oxygen and nitrogen (in air) [1]
 react at high temperatures (and high pressure) [1]
 If nitrogen in fuel [0] out of [2]
- (ii) catalytic converter
 react with carbon monoxide **or** hydrocarbons
 form nitrogen
 ANY TWO [2]
- (d) Add excess lead oxide to nitric acid [1]
 can imply excess
 filter NOT if residue is lead nitrate [1]
 evaporate **or** heat solution [1]

TOTAL = 14

- 5 (a) protons 2
 electrons 2
 neutrons 4 [3]
- (b) (i) $\text{La}^{3+} + 3\text{e}^- = \text{La}$ [1]
 (ii) hydrogen [1]
 bromine NOT Bromide [1]
 caesium hydroxide [1]
 ignore any comments about electrodes



- (c) metal hydroxide or hydroxide ions
hydrogen
- (d) correct formula $1Ba$ to $2Cl$
charges correct
8e around the anion
All three points [2]
Two points ONLY [1]
If covalent [0] out [2]
- (e) alternating (positive and negative) [1]
pattern [1]
- (f) (i) barium - oxygen or ionic [1]
(ii) bond forming energy released/exothermic [1]
bond breaking energy taken in/endergonic [1]
more energy released [1]

TOTAL = 17

Total for Paper: 80

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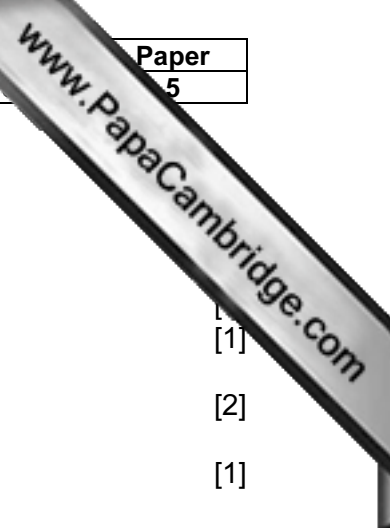
MARKING SCHEME

MAXIMUM MARK: 40

SYLLABUS/COMPONENT: 0620/05

CHEMISTRY

(Practical)



- 1 Table of results
- Experiment 1 Initial and final readings recorded to 1 decimal place
- Experiment 2 Initial and final readings recorded to 1 decimal place [1]
- Results comparable to Supervisor's results $\pm 1 \text{ cm}^3$ [2]
- (a) red/burgundy/brown [1]
- (b) yellow (1) to blue/black (1) see Supervisor IGNORE green [2]
- (c) (i) Experiment 1 [1]
(ii) $\Delta 2 \times$, double volume (1) in Experiment 1 (1) not just more [2]
(iii) potassium iodate less concentrated solution **C** than **B** or vice versa not different concentrations [1]
(iv) $2 \times$ volume from table for Experiment 1 (1) unit (1) [2]
2 x iodine formed [1]
- (d) Indicator (1) reference to accuracy (1)/end-point/see more clearly not test for I_2/I^- [2]

[Question total: 18]

- 2 (a) bubbles/condensation/goes black max 2 [2]
- (b) filtrate - colourless not clear [1]
residue - green [1]
- (c) (i) effervescence/fizz/bubbles [1]
limewater \rightarrow milky [1]
solution is blue [1]
(ii) blue (1) precipitate (1) [2]
royal/deep blue (1) solution (1) [2]
- (d) (i) white (1) precipitate (1) dissolves in excess (1) [3]
(ii) white (1) precipitate (1) dissolves (1) [3]
(iii) white precipitate (1) [1]
- (e) zinc (1) sulphate (1) reversed = 0 [2]
- (f) copper (1) carbonate (1) hydrated (1) reversed = 0 max 2 [2]

[Question total: 22]

[Total for paper: 40]

Results obtained for Question 1/cm³

	1 st	2 nd
Experiment 1	16.5	16.3
Experiment 2	8.3	8.2

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MARKING SCHEME

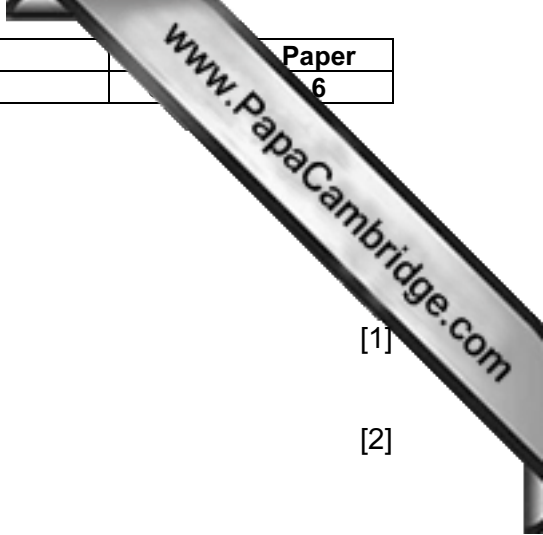
MAXIMUM MARK: 60

SYLLABUS/COMPONENT: 0620/06

CHEMISTRY

(Alternative to Practical)

- 1 (a) A = mortar (1)
 B = stirrer/stirring rod (1) not thermometer
 C = tripod (1)
 D = Bunsen Burner (1) [1]
- (b) filtration [1]
- (c) D or description [1]
- 2 (a) because precipitate formed/goes cloudy (1)
 sulphur (1)/turbid [2]
- (b) reference to fair test/comparison/same depth [1]
- (c) sodium thiosulphate/water 1st/2nd acid, last [1]
- (d) (i) all points correct (3), -1 for any incorrect
 smooth line (1) label (1) [5]
 (ii) line lower down (1)
 does not touch other line (1) [2]
- (e) times would be longer (1) because solution more spread out/reference to
 surface area/depth (1) [2]
- 3 Table of results
 correct burette readings in table (3) or 17.2, 18.9, 26.5
 i.e. 16.8, 17.1 and 25.5
 Differences correctly completed (1) Difference 7.6
 i.e. 8.4 [4]
- (a) (i) Experiment 1 [1]
 (ii) twice volume/more than twice as much [1]
 (iii) Solution B was 2x (1) concentration of C (1) or similar
 B more concentrated than C (1 only) [2]
 (iv) volume A = 33.6 (1) cm³ (1)/34.4cm³
 2x iodine produced (1) [3]
- (b) reference to accuracy (1) indicator (1)/easier to see
not test for I₂ max 2 [2]
- 4 (c) effervescence/fizz/bubbles (1)
 limewater milky (1)/blue solution [2]
- (d) (ii) blue (1) precipitate (1)
 royal/dark blue (1) solution (1) [4]
- (e) (i) white (1) precipitate (1)
 dissolves (1) [3]
 (ii) white (1) precipitate (1)
 dissolves (1) [3]
- (f) Solid D is a sulphate (1) hydrated (1) [2]
- (g) copper (1)/Cu²⁺ (2) [2]



- 5 (a) (i) Smooth line graph
(ii) result at 5 minutes (1)
not on curve (1)/gas escapes, gone down
- (b) 0.8 g [1]
- (c) reference to leak/loss of gas (1)
∴ volumes lower (1) [2]
- 6 Known mass of beach sand (1)
add excess (1) dilute hydrochloric acid (1)
filter (1) wash (1) dry (1) residue
and weigh sand (1) working out result (1)
max 6 of 8 [6]

[Total: 60]

Grade thresholds taken for Syllabus 0620 (Chemistry) in the June 2003 examination

	maximum mark available	minimum mark required for grade:			
		A	C	E	F
Component 1	40	-	26	20	17
Component 2	80	-	52	36	27
Component 3	80	53	31	-	-
Component 5	40	31	24	18	14
Component 6	60	42	32	21	15

The threshold (minimum mark) for B is set halfway between those for Grades A and C.
The threshold (minimum mark) for D is set halfway between those for Grades C and E.
The threshold (minimum mark) for G is set as many marks below the F threshold as the E threshold is above it.
Grade A* does not exist at the level of an individual component.